

Download O2 Equation

Assumptions. The equation relies on the following assumptions: Inspired gas contains no carbon dioxide (CO₂) Nitrogen (and any other gases except oxygen) in the inspired gas are in equilibrium with their dissolved states in the blood. The balanced equation is $2\text{SO}_2 + \text{O}_2 \rightarrow 2\text{SO}_3$. > You follow a systematic procedure to balance the equation. Start with the unbalanced equation: $\text{SO}_2 + \text{O}_2 \rightarrow \text{SO}_3$. A method that often works is to balance everything other than "O" and "H" first, then balance "O", and finally balance "H". Another useful procedure is to start with ...

Answer:- Explanations:- A balanced equation obeys the law of conservation of mass. First equation is not balanced as it has two Fe on left side and four Fe on right side. Below is the balanced equation for chemical compound $\text{S}_8 + 8\text{O}_2 \rightarrow 8\text{SO}_2$. Molar mass $\text{S}_8 = 256.52$ $\text{O}_2 = 31.9988$ $\text{SO}_2 = 64.0638$. A balanced chemical equation happens when the quantity of the particles is required on the reactants side is equivalent to the quantity in the products side.